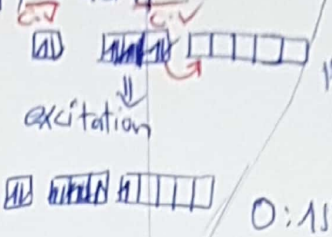
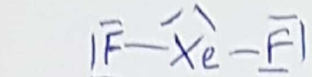
$_{54}Xe : 1s^2/2s^2 2p^6/3s^2 3p^6/4s^2 3d^{10} 4p^6/5s^2 4d^{10} 5p^6 5d^0$

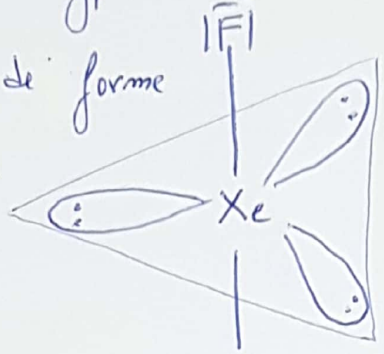


Les trois orbitales sont Hybridés sp^2



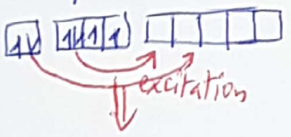
XeF_2 est de type AX_2E_3

XeF_2 est de forme linéaire

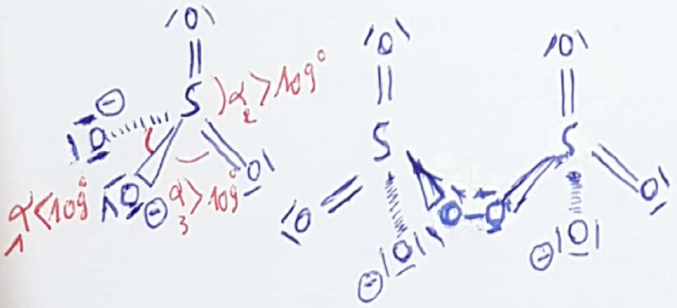


SO_4^{-2} et $S_2O_8^{-2}$

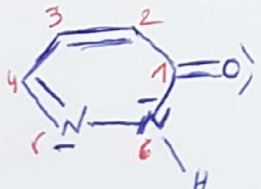
$_{16}S : 1s^2 2s^2 2p^6 / 3s^2 3p^4 3d^0$



SO_4^{-2} est de type AX_4 => Donc sa géométrie est tétraédrique

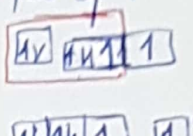


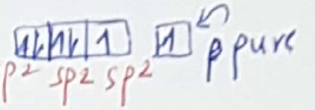
3) La pyridazinone :



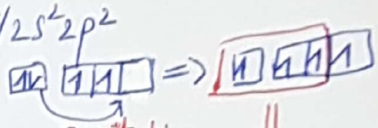
L'hybridation des atomes O, C, N :

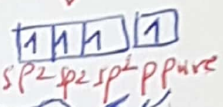
$_8O : 1s^2/2s^2 2p^4$

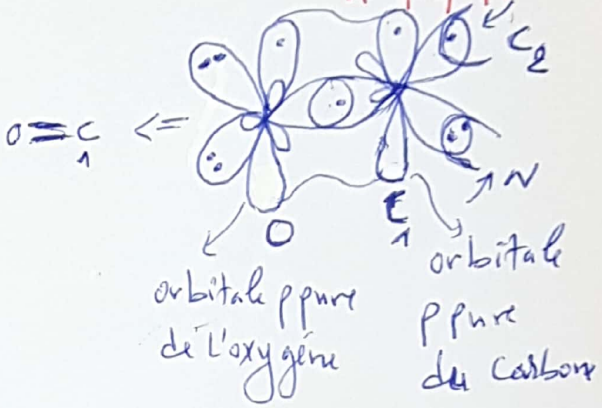




$_6C : 1s^2/2s^2 2p^2$







Hybridation du carbone C_2 est sp^2

le carbone C_3 => sp^2

// C_4 => sp^2

L'azote (N_5) => sp^2

L'azote N_6 => sp^2 malgré il est de type AX_3E par contre le doublet libre de l'azote (N_6) dans la molécule cyclique il n'est pas libre, il est figé dans le cycle donc, le doublet ne participe pas dans la résonance.

$_6C : 1s^2 2s^2 2p^2$ excitation (les 3 orbitales de même forme sont Hybridés sp^2) (3)